** ST. JOSEPH’S COLLEGE (AUTONOMOUS), BANGALORE – 27**

DATE: 8-4-19

**B.Sc CHEMISTRY – VI SEMESTER**

**SEMESTER EXAMINATION- APRIL 2019**

**CH 6115- Inorganic Chemistry**

**SUPPLEMENTARY CANDIDATES ONLY**

**Time : 2 ½ hours Max. Marks: 70**

**This paper contains 2 printed pages and three parts**

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| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Element | Sc | Ti | V | Cr | Mn | Fe | Co | Ni | Cu | Zn |
| Atomic number | 21 | 22 | 23 | 24 | 25 | 26 | 27 | 28 | 29 | 30 |

**Part A**

Answer any **six** of the following. **6 x 2 = 12 M**

1. The outer electronic configuration of a d- block element is 3d104s1. How many ’d’ electrons the metal ion has when it’s oxidation state is +2?
2. What is Ziegler-Natta polymerization?
3. Give one example each for an ambidentate ligand and a bidentate ligand.
4. Hybridisation of central metal ions in two complexes is dsp2 and d2sp3. What will be their geometries?
5. Give two consequences of lanthanide contraction.
6. Write two significant features of Ellingham’s diagram.
7. What is a Pourbaix diagram?
8. Give any two roles of copper in biological systems.

**Part B**

Answer any **eight** of the following  **8 x 6 = 48M**

1. Discuss the following characteristic properties of d- block elements.

i) variable oxidation states ii) magnetic properties.

1. a) Calculate EAN of the central metal ions in the following complexes.

i) Ni[(CO)4] ii) K4[Fe(CN)6]

b) Draw the optical isomers of [Co(en)2Cl2]+. Write the structure of another isomer of this complex which does not exhibit optical isomerism. (3+3)M

11. a) Give the IUPAC names of the following i) [Co(NH3)6 ][Cr(CN)6] ii) Fe(CO)5

iii) [Pt(NH3)5Cl]Br3

b) Write the structural formulae of the following complexes.

i) Hexaaquachromium(III)chloride ii) Tetracyanidonickelate(II) ion

1. Chloridobis(ethylenediamine) nitrito-қOcobalt(III)ion (3+3)M
2. [Co(NH3)6]3+ is diamagnetic whereas [CoF6]3- is paramagnetic. Explain with the help

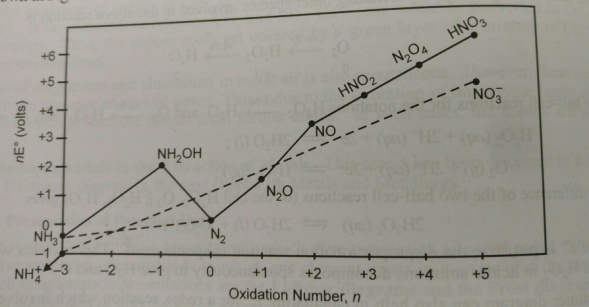
of valence bond theory (VBT).

1. Give one method of preparation and structure of each of the following: i) Mn2(CO)10 ii) W(CH3)6.
2. With an orbital overlap diagram, discuss the bonding in metal carbonyls. What is meant by synergic effect?
3. Using ion exchange chromatography how do you separate individual lanthanides from its solution? Write relevant equations involved in the process.
4. What are metalloporphyrins? Give the structure of a haeme unit. Discuss the oxygen binding characteristics of haemoglobin and myoglobin.
5. How is lithium extracted from lepidolite?
6. One of the complexes of chromium, [CrF6]3- is green in color whereas another chromium complex, [Cr(CN)6]3-, is yellow in color. Explain using spectrochemical series.

**Part C**

Answer any **two** of the following **2 x 5 = 10M**

1. The crystal field stabilisation energy in a tetrahedral complex is found to be 0Dq with no pairing energy. Draw the crystal field splitting pattern with electron arrangement and calculate the magnetic moment of this complex.
2. If ‘M’ is the metal ion with oxidation state +3 and coordination number 6 and if ‘A’ is an anionic uninegative species and ‘B’ is a neutral species, write the formulae of the coordination compounds formed by the combination of these species with the metal ion which are i) most conductive and ii) least conductive.
3. From the Frost diagram for nitrogen given below answer the following i) Identify the most stable species of nitrogen and give reason for the choice ii) Identify the species which undergoes disproportionation and give the corresponding equation iii) list three species which are equally stable.



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